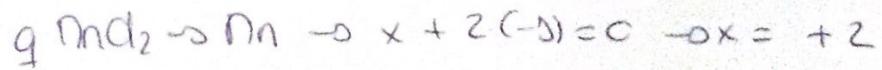
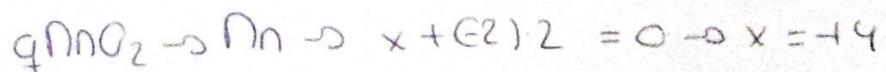
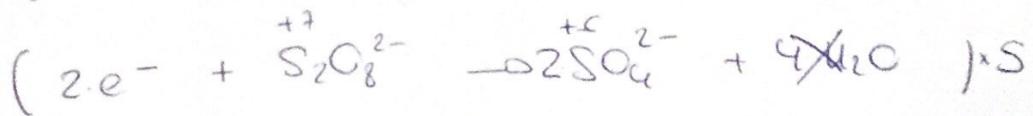
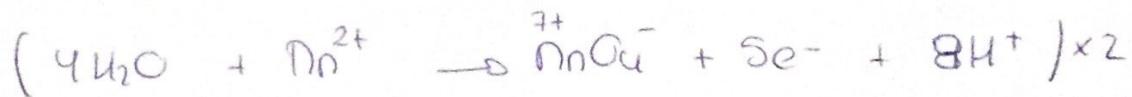
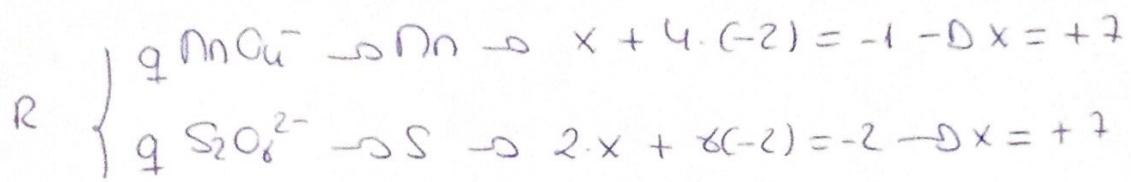
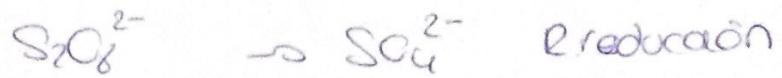
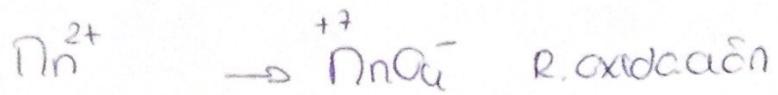
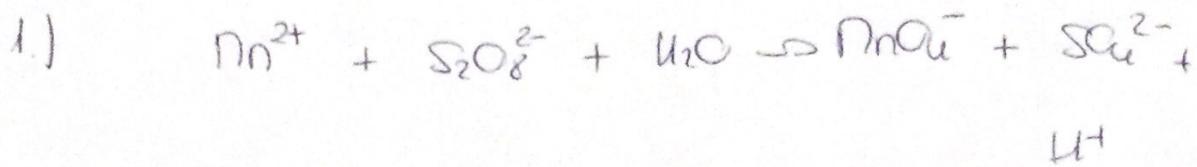
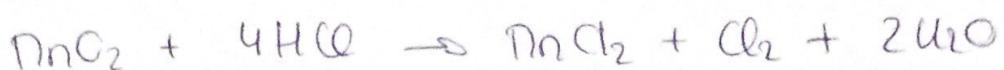
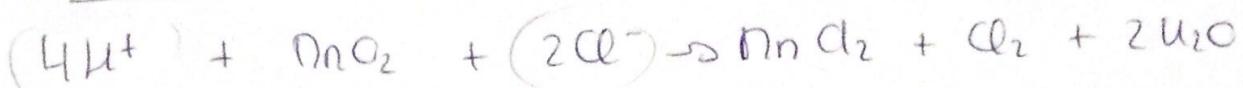
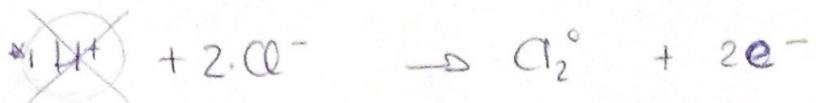
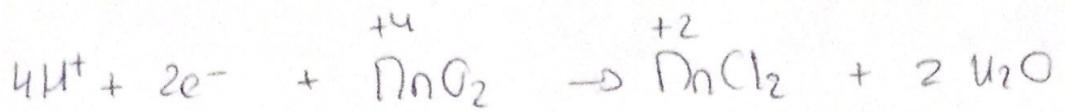


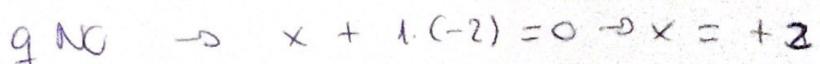
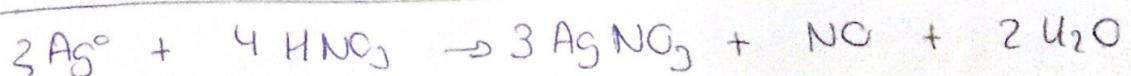
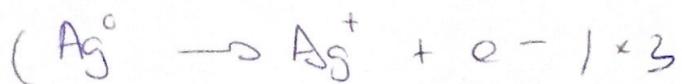
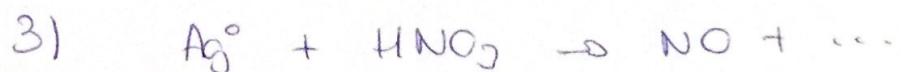
BLOQUE III: EQUILIBRIOS

1



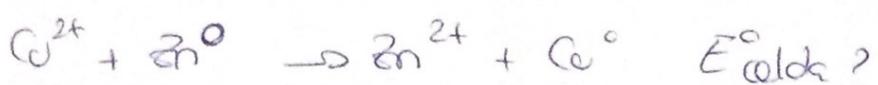
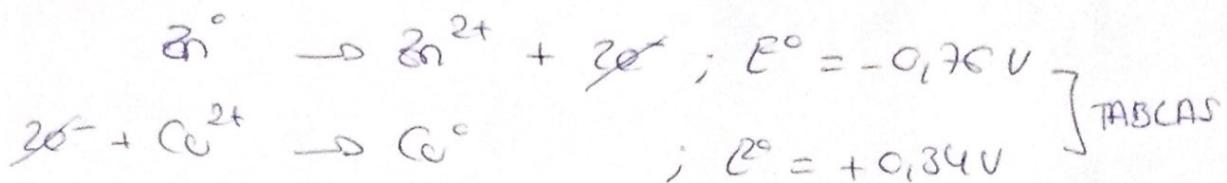


Q. → es el medio ácido de la propia reacción → ya lo ajustamos con el $\text{DnO}_2 / \text{DnCl}_2$. El H_2O "controla" la acidez del medio.



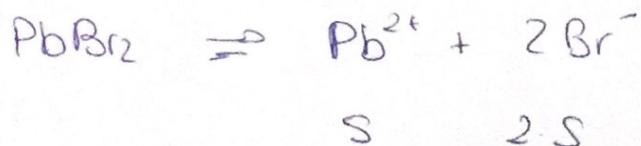
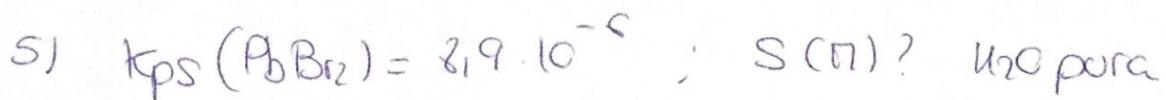
2

4) E° ? $\text{Br}|\text{Br}^{2+} // \text{Cl}^{2+}|\text{Cl}$



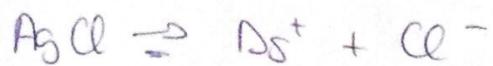
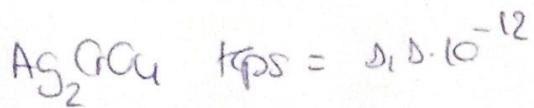
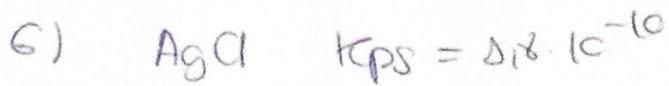
$$E_{\text{cold}}^c = +0.34(\nu) - (-0.75(\nu))$$

E_c elda = 1,1 (v) ESPONTÂNEA

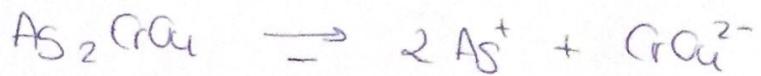


$$t_{\text{GPS}} = (\text{Pb}^{2+}) (\text{Br}^{2-})^2 = s \cdot (2.8)^2 = 8,9 \cdot 10^{-6}$$

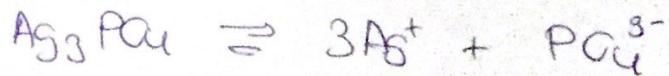
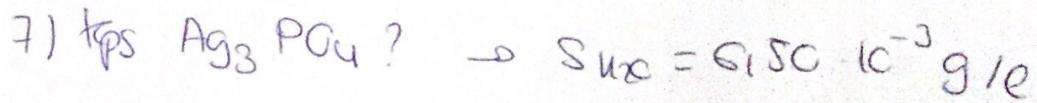
$$S = 010130\pi$$



$$k_{\text{PS}} = 1,8 \cdot 10^{-10} = S \cdot S \rightarrow S = 1,3 \cdot 10^{-5} \text{ M}$$



$$k_{\text{PS}} = 1,1 \cdot 10^{-12} = (2S)^2 \cdot S \rightarrow S = 6,5 \cdot 10^{-5} \text{ M}$$



$$k_{\text{PS}} = (3 \cdot S)^3 \cdot S \rightarrow k_{\text{PS}} = (3 \cdot 6,15 \cdot 10^{-3})^3 \cdot 6,15 \cdot 10^{-3}$$

$$k_{\text{PS}} = 1,56 \cdot 10^{-18}$$

/