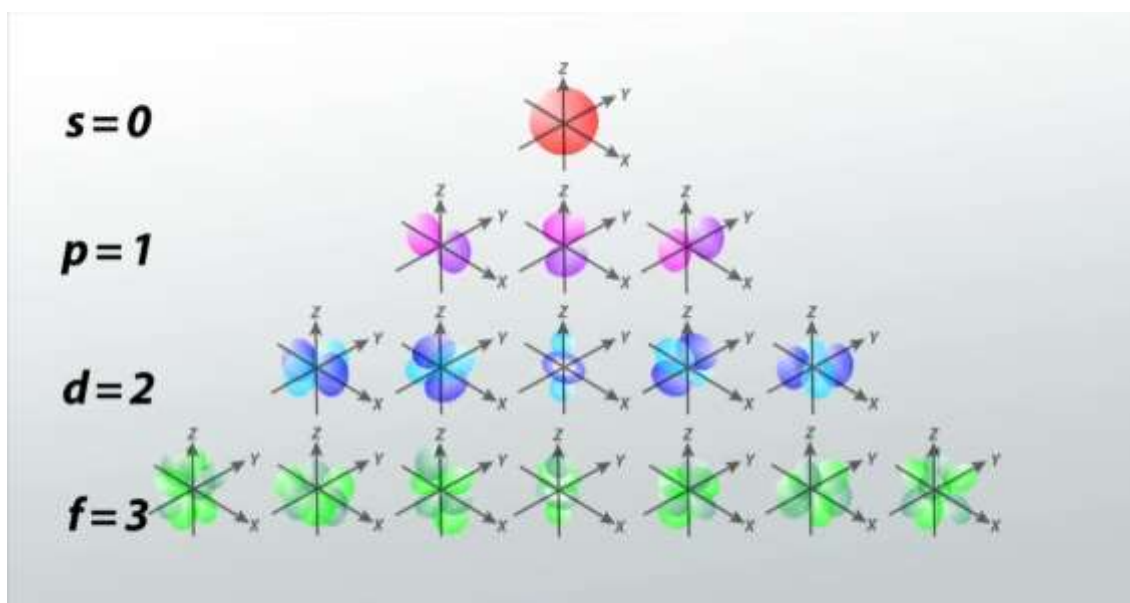


NUMEROS CUANTICOS

Subnivel (l)	Orbitales	Número de orbitales	Número máximo de electrones
s (l = 0)	$\frac{\uparrow\downarrow}{0}$	1	2
p (l = 1)	$\frac{\uparrow\downarrow \quad \uparrow\downarrow \quad \uparrow\downarrow}{-1 \quad 0 \quad +1}$	3	6
d (l = 2)	$\frac{\uparrow\downarrow \quad \uparrow\downarrow \quad \uparrow\downarrow \quad \uparrow\downarrow \quad \uparrow\downarrow}{-2 \quad -1 \quad 0 \quad +1 \quad +2}$	5	10
f (l = 3)	$\frac{\uparrow\downarrow \quad \uparrow\downarrow \quad \uparrow\downarrow \quad \uparrow\downarrow \quad \uparrow\downarrow \quad \uparrow\downarrow \quad \uparrow\downarrow}{-3 \quad -2 \quad -1 \quad 0 \quad +1 \quad +2 \quad +3}$	7	14



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