UNIVERSITY CEU SAN PABLO
SCHOOL OF PHARMACY
DEPARTMENT OF CHEMISTRY AND BIOCHEMISTRY

PROBLEMS OF PHYSICAL CHEMISTRY

2018-2019

LESSON 3

5. Calculate the change in pressure required to reduce the normal freezing temperature by 1 degree for:
a) water
b) a solvent A
Data: $R = 0.082\ l\cdot atm\cdot K^{-1}\cdot mol^{-1} = 1.987\ cal\cdot K^{-1}\cdot mol^{-1} = 8.314\ J\cdot K^{-1}\cdot mol^{-1}$
$\Delta\overline{H}_{m,A} = 97.7\ J\cdot mol^{-1}$; $\Delta\overline{H}_{m,H2O} = 79.7\ cal\cdot g^{-1}$; $T^{o}_{m,A} = 156\ K$; $M_A = 74.12\ g\cdot mol^{-1}$;
$\rho\ (H_2O_{(l)}) = 0.9998\ g\cdot ml^{-1}$; $\rho\ (H_2O_{(s)}) = 0.9168\ g\cdot ml^{-1}$; $\rho\ (A_{(l)}) = 0.715\ g\cdot ml^{-1}$;
$\rho\ (A_{(s)}) = 0.820\ g\cdot ml^{-1}$

6. The vapour pressures of solid and liquid HCN can be expressed by the following equations:

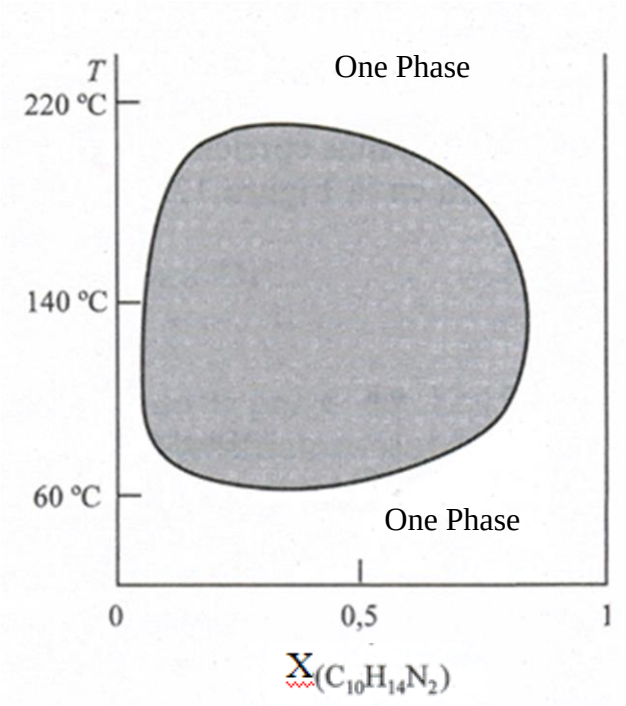
$$\text{Solid}: \quad \text{Log}_{10}\,P(\text{mmHg}) = 9.339 - \frac{1865}{T}(K)$$

$$\text{Liquid}: \quad \text{Log}_{10}\,P(\text{mmHg}) = 7.745 - \frac{1453}{T}(K)$$

Determine:
a) the normal boiling point and the triple point
b) the sublimation, vaporization and melting heat of HCN
Data: $R = 0.082\ l\cdot atm\cdot K^{-1}\cdot mol^{-1} = 1.987\ cal\cdot K^{-1}\cdot mol^{-1} = 8.314\ J\cdot K^{-1}\cdot mol^{-1}$

the masses of water and nicotine change if temperature is raised to 140 ° C? And if pressure is changed?

Data: M ($C_{10}H_{14}N_2$) = 162 g·mol^{-1}; M (H_2O) = 18 g·mol^{-1}



9.- According to the following phase diagram, obtained for naphthalene and p-nitrotoluene, indicate:

a) The melting temperature of the two pure solids.

b) What represent the AB and BC curves and the B point?

c) What is the temperature when naphthalene starts to freeze in the cooling process of a mixture with $x_{p\text{-nitrotoluene}}$ = 0.20?

d) If the same solution is cooled to 40 ºC, what phases are present? What are their compositions? What is the relative amount?