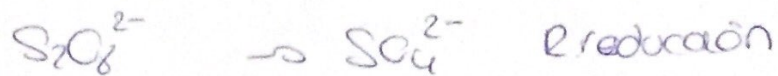
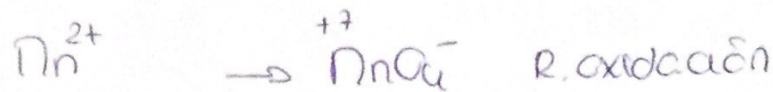
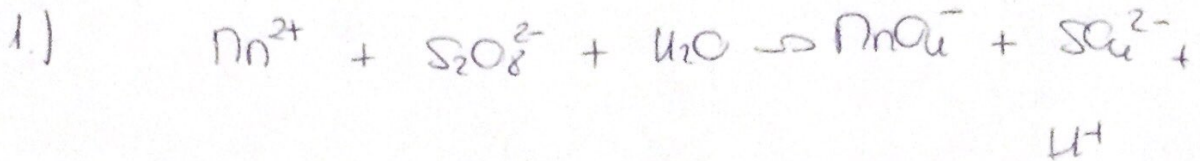


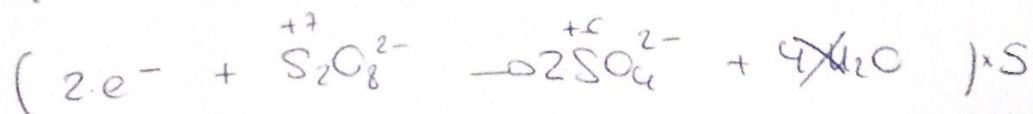
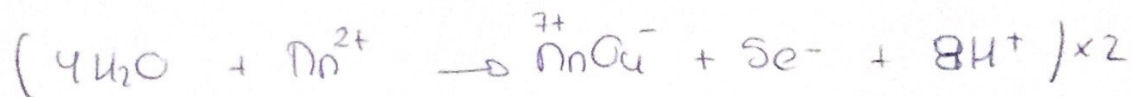
BLOQUE III: EQUILIBRIOS

1



$$\text{R} \left\{ \begin{array}{l} q \text{ MnO}_4^- \rightarrow \text{Mn} \rightarrow x + 4(-2) = -1 \rightarrow x = +7 \\ q \text{ S}_2\text{O}_8^{2-} \rightarrow \text{S} \rightarrow 2x + 8(-2) = -2 \rightarrow x = +7 \end{array} \right.$$

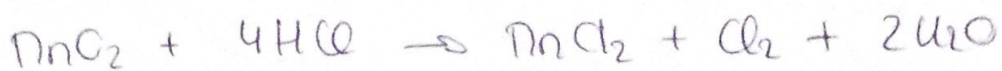
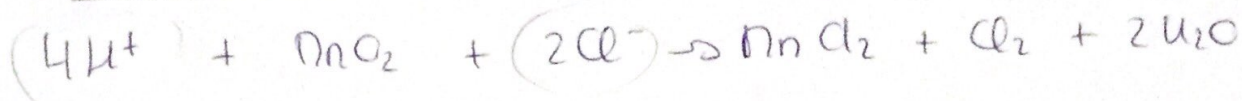
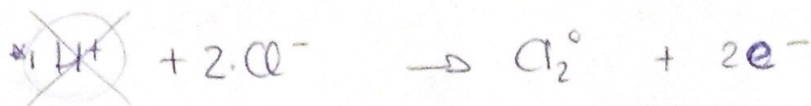
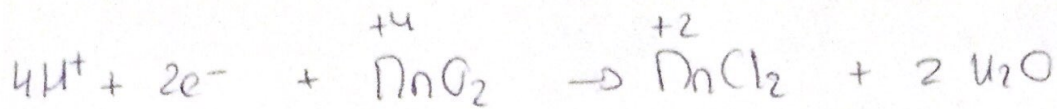
$$\text{P} \left\{ \begin{array}{l} q \text{ SO}_4^{2-} \rightarrow \text{S} \rightarrow x + 4(-2) = -2 \rightarrow x = +6 \end{array} \right.$$



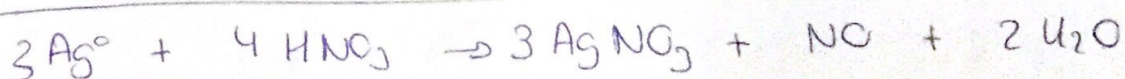
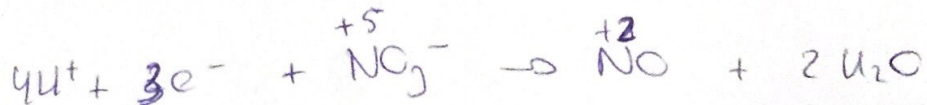
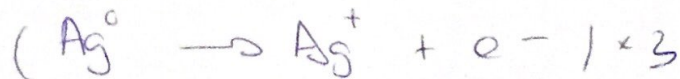
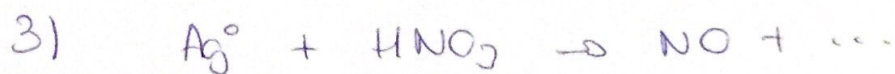
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②, → es el medio ácido de la propia reacción → y se le sistemas con el $\text{MnO}_2 / \text{MnCl}_2$. El H_2O "controla" la acidez del medio.



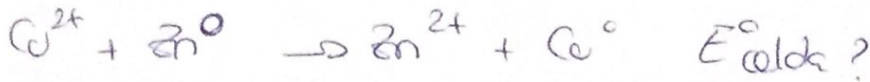
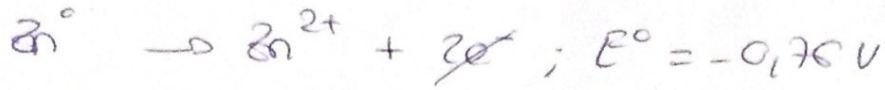
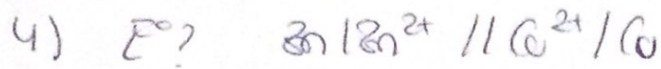
$$9 \text{ NO}_3^- \rightarrow x + 3(-2) = -1 \rightarrow x = +5$$

$$9 \text{ NO} \rightarrow x + 1(-2) = 0 \rightarrow x = +2$$

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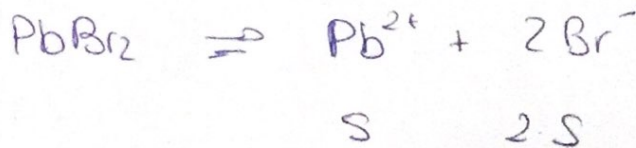
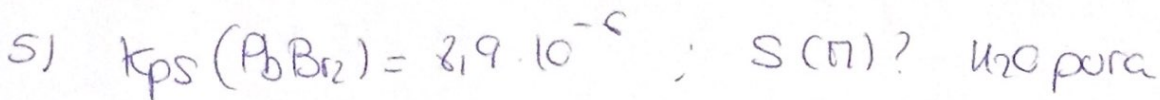
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E° celda = E° CATODO (R. Reducción) - E° ANODO (R. oxidación)

E° celda = +0,34(V) - (-0,76(V))

E° celda = 1,1 (V) ESPONTÁNEA



Kps = (Pb²⁺)(Br²⁻)² = S · (2·S)² = 8,9 · 10⁻⁶

S = 0,01307

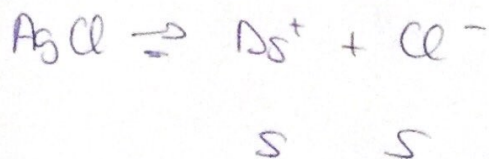


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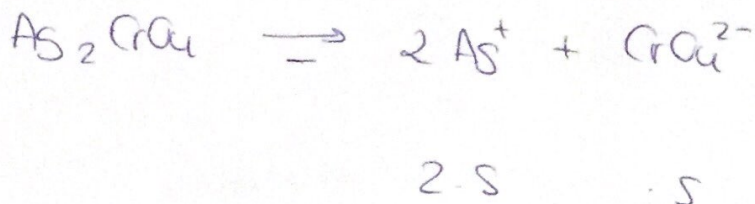
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$$6) \text{AgCl} \quad K_{ps} = 1,8 \cdot 10^{-10}$$

$$\text{Ag}_2\text{CrO}_4 \quad K_{ps} = 1,1 \cdot 10^{-12}$$

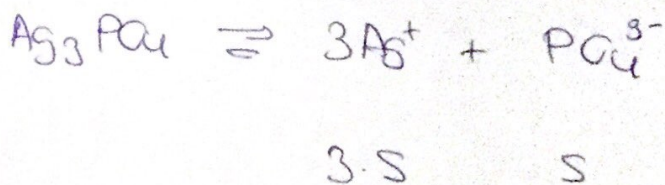


$$K_{ps} = 1,8 \cdot 10^{-10} = S \cdot S \rightarrow S = 1,3 \cdot 10^{-5} \text{ M}$$



$$K_{ps} = 1,1 \cdot 10^{-12} = (2 \cdot S)^2 \cdot S \rightarrow S = 6,5 \cdot 10^{-5} \text{ M}$$

$$7) K_{ps} \text{Ag}_3\text{PO}_4? \rightarrow S_{\text{Ag}_3\text{PO}_4} = 6,50 \cdot 10^{-3} \text{ g/l}$$



$$K_{ps} = (3 \cdot S)^3 \cdot S = 27 \cdot S^4$$

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